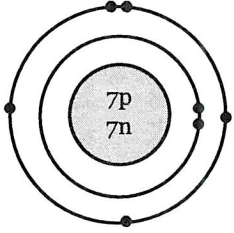
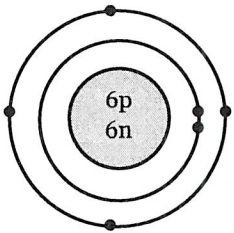
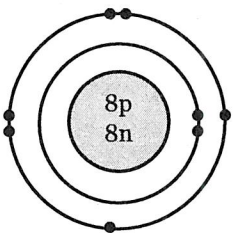
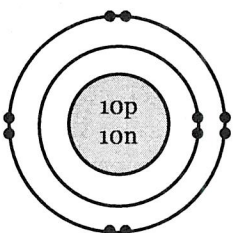


Use your periodic table to answer the following.

	a. number of protons <u>7</u>
	b. number of electron shells <u>2</u>
	c. number of electrons <u>7</u>
	d. number of electrons in outer shell <u>5</u>
	e. element <u>nitrogen (N)</u>
	a. number of protons <u>6</u>
	b. number of electron shells <u>2</u>
	c. number of electrons <u>6</u>
	d. number of electrons in outer shell <u>4</u>
	e. element <u>carbon (C)</u>
	a. number of protons <u>8</u>
	b. number of electron shells <u>2</u>
	c. number of electrons <u>8</u>
	d. number of electrons in outer shell <u>6</u>
	e. element <u>oxygen (O)</u>
	a. number of protons <u>10</u>
	b. number of electron shells <u>2</u>
	c. number of electrons <u>10</u>
	d. number of electrons in outer shell <u>8</u>
	e. element <u>neon (Ne)</u>

These four elements are all in the same horizontal row (period) of the periodic table. What is the **same** about electron shells for elements in the same period?

Elements in the same period have the same number of energy shells occupied by electrons.

What is **different** about the electrons in the outer shell for elements in the same period?

They have different numbers of electrons in the valence shell.  
(They also often have different numbers of protons, neutrons.)