

Homework

ASSIGNMENT #1: Bohr Model Practice, Worksheet pages 6-8
 This assignment is to be completed below in the space provided.

Use the innermost circle as the nucleus, and fill the electron shells with the correct number of electrons for each of the first 20 elements in the Periodic Table.
 eg. Hydrogen has been completed for you as an example.

Bohr Diagrams for the first 20 Elements

	Group 1	Group 2	Group 3	Group 4	Group 5	Group 6	Group 7	Group 8	
Period 1	H 							He 	
Period 2	Li 	Be 	B 	C 	N 	O 	F 	Ne 	
Period 3	Na 	Mg 	Al 	Si 	P 	S 	Cl 	Ar 	
Period 4	K 	Ca 							

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1. What is the pattern between the number of **valence electrons** and the group number of the periodic table?

The "ones place" of the group number is the same as the number of valence electrons.

Except: hydrogen has 1 valence electron
 helium has 2 valence electrons.

2. What is the pattern between the number of **electron shells** and the period number of the periodic table?

Period number = # of energy shells with electrons