

Use with textbook pages 186–196.

Chemical names and formulas of ionic compounds

1. Write the name for each of the following compounds.

(a) BeS

(b) Hg₃N₂

(c) Cu(NO₃)₂

(d) Ag₂O

(e) CoBr₂

(f) Bi₃(PO₄)₅

(g) CaF₂

(h) Mn₂O₃

(i) Cr₂(SO₄)₃

(j) ZnCl₂

(k) Ni(OH)₂

(l) K₂Cr₂O₇

(m) ScF₃

(n) NaI

(o) Pb(CO₃)₂

(p) RbClO₂

(q) K₃P

(r) Mg(CN)₂

(s) SnS

(t) NaHCO₃

2. Write the chemical formula for each of the following compounds.

(a) aluminum bromide _____

(b) platinum(II) sulphide _____

(c) strontium sulfite _____

(d) scandium oxide _____

(e) titanium(IV) nitrite _____

(f) ammonium sulphate _____

(g) lithium selenide _____

(h) lead(II) hydrogen sulphate _____

(i) sodium acetate _____

(j) cesium chloride _____

(k) cadmium(II) hydroxide _____

(l) zinc phosphate _____

(m) barium chloride _____

(n) tin(II) permanganate _____

(o) lithium hypochlorite _____

(p) gold(III) sulphate _____

(q) sodium nitrate _____

(r) chromium(III) chloride _____

(s) potassium carbonate _____

(t) iron(III) bisulphate _____

Use with textbook pages 193–197.

Chemical names and formulas of covalent compounds

1. What is a covalent compound?

2. What type of bond is formed in a covalent compound?

3. What is used in naming covalent compounds?

4. Write the chemical formula for each of the following compounds.

(a) silicon dioxide _____	(i) dinitrogen pentoxide _____
(b) chlorine dioxide _____	(j) dinitrogen monoxide _____
(c) tellurium dioxide _____	(k) arsenic tetrabromide _____
(d) selenium trioxide _____	(l) arsenic pentachloride _____
(e) carbon disulphide _____	(m) disulphide pentoxide _____
(f) arsenic trichloride _____	(n) sulphur monochloride _____
(g) chlorine heptoxide _____	(o) phosphorus trichloride _____
(h) selenium difluoride _____	(p) diphosphorus pentoxide _____