

Naming Compounds & Counting Atoms

.....A Lesson in Names & Numbers

Try This!

Determine the chemical formulas for the following:



Calcium + Chlorine		Silver + Hydroxide		Copper (II) + Nitrogen	
Write Ions Here	Write Formula Here	Write Ions Here	Write Formula Here	Write Ions Here	Write Formula Here

What is in a Name?.....

Naming Metals (with one ion charge)

- The name of the metal ion is written _____
- The name is written in _____ and spelled exactly the same as the element name
ex. Al is _____

Naming Non-Metal Ions

- The name of the non-metal ion is written _____, after the metal.
- The name is written almost the same the element name except the ending is changed to _____ to distinguish from Polyatomic Ions
ex. Cl is _____ and O is _____

Naming Polyatomic Ions

- Positive polyatomic ions are written _____. There is only one, which is _____.
- Negative polyatomic ions are written _____ and the name of the ion is not changed.
ex. SO_4^{+2} is _____

Naming Multivalent Metals

- If the metal is multivalent, like Iron (Fe^{+2} and Fe^{+3}), the ion charge of the metal must be in the name. This charge is indicated by _____.
ex. Cu^{+2} is _____ and read as “_____”

POSITIVE (usually metal) ION FIRST	NEGATIVE (non-metal) ION SECOND

Put it into Practice

	IONS	NAME OF COMPOUND	FORMULA
1	K ⁺¹ Cl ⁻¹		
2	Na ⁺¹ N ⁻³		
3	Ca OH		
4		aluminum hydroxide	
5	K N		
6		calcium oxide	
7		sodium chloride	
8	Fe ⁺² O		

