NaminG Compounds & CountinG atoms							
Try This!  A Lesson in Names & Number  So you don't have a name?  I haven't been discovered yet, so I'm still unnamed!  I'm still unnamed!  I'm still unnamed!							
Dete	mine the o	chemical formula	s for the followin	g:	(nope!)	they don't call me.	
Writ	Calcium e lons Here	+ Chlorine Write Formula Here	Silver + F Write Ions Here	Hydroxide Write Formula Here	Copper (II) Write Ions Here	+ Nitrogen  Write Formula Here	
W.h.	at is iī	a <b>N</b> am <i>e</i> ?	••••••		•••••	•••••	
Naming Metals (with one ion charge)  The name of the metal ion is written and spelled exactly the same as the element name ex. Al is							
Naming Non-Metal lons  The name of the non-metal ion is written, after the metal.  The name is written almost the same the element name except the ending is changed to to distinguish from Polyatomic lons ex. Cl is and O is							
<ul> <li>Naming Polyatomic lons</li> <li>Positive polyatomic ions are written There is only one, which is</li> <li>Negative polyatomic ions are written and the name of the ion is not changed.</li> <li>ex. SO<sub>4</sub><sup>+2</sup> is</li> </ul>							
		alent Metals	libe Iron (Fe <sup>+2</sup> ar	nd Fe <sup>+3</sup> ), the ion cl	harae of the met	al must be in	

the name. This charge is indicated by \_\_\_\_\_ and read as "\_\_\_

## []he □UmmaRy

POSITIVE (c	usually metal) ION FIR\$T	NEGATIVE (non-metal) ION \$ECOND

## Put it into Practice

	IONS	NAME OF COMPOUND	FORMULA
1	K <sup>+1</sup> Cl <sup>-1</sup>		
2	Na <sup>+1</sup> N <sup>-3</sup>		
3	Са ОН		
4		aluminum hydroxide	
5	K N		
6		calcium oxide	
7		sodium chloride	
8	Fe <sup>+2</sup> O		

9		iron (III) oxide					
10		magnesium phosphate					
11	Fe (III) Cl						
12		potassium phosphate					
13		hydrogen hydroxide					
14	Са СОз	calcium carbonate					
The Last Bit  Counting Atoms 1234551890123455189							
A small number called a, right next to a letter, indicates the number of atoms of an element in that compound.							
If the subscript is next to a, then it multiples every element inside the bracket by that number.							
How many atoms of each element are in the compounds below?							
	Fe <sub>2</sub> O <sub>3</sub>		Fe O				

Fe(OH)<sub>2</sub> FeOH<sub>2</sub>